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Match the type of mixture on the left to its representative particle diameter on the right.

c solutions (a) larger than 1000 nm a

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suspensions (b) 1 nm to 1000 nm b colloids  
(c) smaller than 1 nm 2.

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suspensions (b) 1 nm to 1000 nm b colloids  
(c) smaller than 1 nm 2. Identify the solvent  
in each of the following examples:

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Answers Chapter 12 SECTION 1 SHORT

ANSWER 1. c 2. a 3. b 2. a. alcohol b. water

c. the gels 3. The mixture is a colloid. The properties are consistent with those reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture

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Exercise 13D; Exercise 13E; Exercise 13F;  
Exercise 13G; Mixed Exercise 13; Ch. 14:  
Exponentials and Logarithms Exercise 14A  
Exercise 14B Exercise 14C Exercise 14D  
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Answers Chapter 12 SECTION 1 SHORT

ANSWER 1. c 2. a 3. b 2. a. alcohol b. water

c. the gels 3. The mixture is a colloid. The

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properties are consistent with those reported in Table 3 on page 404 of the text.

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following questions in the space provided. 1.  
Solid  $\text{CaCl}_2$  does not conduct electricity.  
Explain why it is considered to be an  
electrolyte.  $\text{CaCl}_2$  is an ionic solid. In the  
crystal form, its ions are locked in position.  
Dissolving

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